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1754 'base' +
acid→solid 1832 role of
hydrogen: acids
contain hydrogen, can
be replaced by a metal

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1884 Arrhenius model:
acids release H^+ ,
bases release OH^- ions
(in aq solution) 1923
Bronsted-Lowry model
(central role protons)
Bronsted Lowry
acid=proton donor,
Bronsted-Lowry
base=proton acceptor

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20.1 DESCRIBING ACIDS AND BASES

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HYDROGEN IONS AND ACIDITY FEATURES

These tart-tasting fruits contain a weak acid called citric acid.

DISCOVER IT! Effect of Foods on Baking Soda

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water and salt; 4
Formulas of Acids

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Chapter 20 acids and bases. the particle formed when a base gains a hydrogen ion: NH_4^+ is the conjugate acid of the base

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NH₃*20.3. two substances that are related by the loss or gain of a single hydrogen ion.

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Best indicator for a strong acid-strong base titration: 15. H₂CO₃ is the formula of _____
_____ 20. The higher the value of pK_a the _____ the acid: 21. The number of decimal

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places given for pH values: 22. The K_a of a weak acid can be determined by making a _____ solution of the acid and then measuring the pH with a pH probe.

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particle formed when a
base gains a hydrogen
ion: NH_4^+ is the
conjugate acid of the
base NH_3 *20.3. two
substances that are
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gain of a single
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In this chapter, students will learn about the nature and behaviour of Acids, bases and salts. This chapter mainly describes the chemical nature of acids, bases and salts, their reaction with metals, non-metals and with each other. This chapter is quite interesting as it involves many practical experiments to help you learn the basics.

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