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Group . 1A : 2A . 3A : 4A . 5A : 6A . 7A : 8A . 1 : 2 . 3 : 4 • Notice that all the electrons within a given group (with the exception of helium) have the same ...

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## **Chemistry Ionic And Metallic Bonding Answers**

Metallic Bonds and Metallic Properties 1. Is the following sentence true or false? Metals are made up of cations and valence electrons, not neutral atoms. 2. What are metallic bonds? 3. Name three properties of metals that can be explained by metallic bonding. a. b. c. 4. What happens to an ionic crystal when a force is applied to it? 5.

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## **BONDING AND INTERACTIONS**

When the elements react bonds are formed. Types of Bonds  
Ionic Metallic Covalent Ionic Bonds Bonds that are formed by transfer of electrons from one element to the other. Each element (now an ion) will have a complete octet after the transfer of electrons. The Ionic Bond • The electrical force of attraction between oppositely charged ions.

## **PowerPoint Presentation**

The key difference between ionic bonding and metallic bonding is that the ionic bonding takes place between positive and negative ions whereas the metallic bonding takes place between positive ions and electrons. As American chemist G.N.Lewis proposed, atoms are stable when they contain eight electrons in their valence shell.

## **Difference Between Ionic Bonding and Metallic Bonding**

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Title: PowerPoint Presentation Author: Debbie Munson Created Date: 1/20/2016 9:08:37 PM

## **Ionic and Metallic Bonding - Pittsfield**

Metallic Bonds and Metallic Properties When a metal is subjected to pressure, the metal cations easily slide past one another. • If an ionic crystal is struck with a hammer, the blow tends to push the positive ions close together. • The positive ions repel one another, and the crystal shatters.

## **Chapter 7.3 Slides**

Definition: A metallic bond is formed when the valence electrons are not associated with a particular atom or ion, but exist as a "cloud" of electrons around the ion centers. Metallic materials have good electrical and thermal conductivity when compared to materials with covalent or ionic bonding. A metal such as iron has metallic bonding.

## **Atomic Bonding - University of Washington**

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When dissolved in water, ionic compounds can conduct electricity. b. When melted, ionic compounds do not conduct electricity. c. Ionic compounds have very unstable structures. d. Ionic compounds are electrically neutral. SECTION 7.3 BONDING IN METALS (pages 201–203) This section uses the theory of metallic bonds to explain the physical proper-

### **SECTION 7.1 IONS (pages 187–193)**

Chapter 7 Ionic and Metallic Bonding 159 \_\_\_\_ 12. The coordination number of an ion is the number of ions of positive charge that surround the ion in a crystal. \_\_\_\_ 13. The coordination number of the ion  $\text{Na}^+$  in  $\text{NaCl}$  is 6. \_\_\_\_ 14. In forming an ionic compound, an atom of an element gains electrons. \_\_\_\_ 15.

### **05 CTR ch07 7/9/04 3:27 PM Page 155 IONS 7**

The Ionic and Metallic Bonding chapter of this Prentice Hall Chemistry Companion Course helps students learn the essential lessons associated with ionic and metallic bonding. Each of these simple...

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